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The most fun you can have with a calculator. More Exciting

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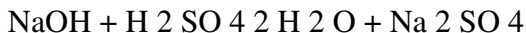
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Balancing Equations and Simple Stoichiometry-KEY Balance the following equations: 1) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$ 2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$ 3) $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + 1 \text{ CO}_2$ 4) $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$ 5) $3 \text{ SnO} + 2 \text{ NF}_3 \rightarrow 3 \text{ SnF}_2 + 1 \text{ N}_2\text{O}_3$ Using the following equation: 2

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Balancing Equations and Simple Stoichiometry-KEY

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mass B the balanced equation Double lined boxes are Conversion
Factors to convert from one quantity to another. mole ratios reactant
s and s y lar mass

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Step 8: Finding our answer. To find the answer to this calculation,
multiply all the terms on the top together ($17.5 \times 1 \times 2 \times 22.4$) and
divide by the product of the terms on the bottom ($28.0 \times 1 \times 1$). If
you do the calculation accurately, you should find that you have
28.0 liters of ammonia gas. Which is the answer.

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Chemfiesta Balancing Equations Worksheet Part 2 Answers
Tessshlo. Balancing Chemical Equations Worksheets With Answers
Science Adding Fractions Test Blank 10 7th Grade Math Counting
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Stoichiometry is a set of calculations you perform to figure out how much stuff you can make in a reaction, or how much stuff you will need to make the reaction occur. In other words, stoichiometry is used to figure out if you've got enough crackers to make 30 sandwiches, or how much cheese you'll need to make 15 sandwiches.

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Updated 2-10-16. Types of reaction and reaction products: Overview of chemical reactions: Some info from a book I never finished about chemical reactions. Pretty much covers the stuff above. Types of reaction: There are several categories into which chemical reactions can be placed, and knowing them can help you to figure out how to predict the ways in which chemical reactions will operate.